2.B.2 - Nitric Acid Production

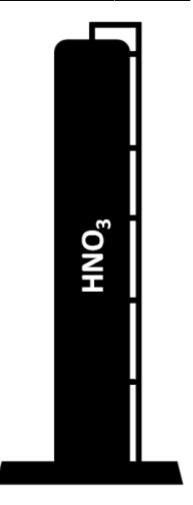
Short description

Category Code		Method					AD					EF						
2.B.2					PS					D						1		
Key Category	SO2	NO×	NH₃	N	чиос	СО	BC	Pb	Hg	Cd	Diox	PAH	НСВ	TSP	PM	10 F	PM2 5	
2.B.2	-	L/-	-		-	-	-	-	-	-	-	-	-	-	-		-]
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Methods																		
	D					Def	ault											
	RA					Ref	eren	ice /	Арр	roa	ch							
	T1					Tier	1/	Sim	ple	Ме	thodo	ology	*					
	Т2					Tier	2*											
	Т3					Tier	3/	Det	aile	d M	etho	dolog	у *					
	С					COF	RINA	IR										
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* as described ir	n the	EME	P/C0	RIN	IAIR Ei	niss	ion	Inve	ento	ory (Guide	book	- 200	7, in	the g	gro	oup s	pecific chapt
AD - Data Soui	ce f	or Ao	tivi	ty I	Data													
NS National Stat	tistic	s																
RS Regional Sta	tistic	S																
IS International	Stat	istics	;															
PS Plant Specifi	c dat	a																
AS Associations	, bus	iness	orga	anis	sation	5												
Q specific ques	stion	naire	s, su	rve	ys													
EF - Emission I	acto	ors																
Default (EME	P Gu	idebo	ook)															
C Confidential																		
CS Country Spee	cific																	
PS Plant Specific	: dat	а																

During the production of nitric acid, nitrogen oxide occurs in a secondary reaction. In Germany, there are currently seven nitric acid production plants. HNO₃, production occurs in two process stages:

- Oxidation of $\ensuremath{\mathsf{NH}}_{\ensuremath{\scriptscriptstyle 3}}$ to NO and
- Conversion of NO to NO₂ and absorption in H₂O.

Details of the process are outlined below:



Catalytic oxidation of ammonia

A mixture of ammonia and air at a ratio of 1:9 is oxidised, in the presence of a platinum catalyst alloyed with rhodium and/or palladium, at a temperature of between 800 and 950 °C. The related reaction, according to the Oswald process, is as follows:

 $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$

Simultaneously, nitrogen, nitrous oxide and water are formed by the following undesired secondary reactions:

4 NH₃ + 3 O₂ -> 2 N₂ + 6 H₂O 4 NH₃ + 4 O₂ -> 2 N₂O + 6 H₂O

All three oxidation reactions are exothermic. Heat may be recovered to produce steam for the process and for export to other plants and/or to preheat the residual gas. The reaction water is condensed in a cooling condenser, during the cooling of the reaction gases, and is then conveyed into the absorption column.

Methodology

Activity data

As this source category is a key category for N_2O , plant specific activity data are applied here according to the IPCC guidelines.

These data are made available basically via a co-operation agreement with the nitric acid producers and the IVA (Industrieverband Agrar). As the data provided by the producers has to be treated as confidential, it is anonymised by the IVA before submitting it to the UBA. Here, one producer is delivering its data directly to the UBA. After checking these specific data, they are merged with the IVA.

According to the IVA, catalytic reduction is used as a reduction method in some of the plants.

Emission factors

For NO_x the default emission factor from the EEA Emission Inventory Guidebook of 10 kg/t $NH_3^{(1)}$ is used.

Recalculations

With **activity data** and **emission factors** remaining unrevised, no recalculations have been carried out compared to last year's submission.



For pollutant-specific information on recalculated emission estimates for Base Year and 2018, please see the pollutant specific recalculation tables following chapter 8.1 - Recalculations.

Planned improvements

Germany is in contact with the IVA with the aim to get a Country-specific EF for the NOx emissions. It is expected that the Country-specific EF will be lower than the Default EF.

¹⁾ EEA, 2013: EMEP EEA Emission Inventory Guidebook 2013, Aug 2013: page 15, table 3.3 Tier 1 emission factor for source category 2.B.2 Nitric acid production